31. The speed of a homogeneous reaction is determined by the equation:

A) $V = \pm\frac{1}{s}\frac{dn_A}{d\tau}$;

B) $V = \pm\frac{1}{s}\frac{dc}{dn_A}$;

C) $V = R \cdot F \cdot C$;

D) $V = k \cdot C_A \cdot C_B$;

E) $V = k(C_A - C_B)^n$.

32. The speed of a homogeneous reaction is determined by the equation:

A) $V = \pm\frac{1}{s}\frac{dn_A}{d\tau}$;

B) $V = \pm\frac{1}{\tau}\frac{dnj}{dc}$;

C) $V = \pm\frac{1}{s}\frac{dc}{dn_A}$;

D) $V = R \cdot F \cdot C$;

E) $V = k \cdot C^{n_{AA}} ¿ C^{n_{\cdot}}$.

33. Ways to intensify homogeneous processes:

A) increase in concentration, decrease in temperature;
B) increase in pressure and concentration;
C) pressure and temperature reduction;
D) decrease in concentration and temperature;
E) pressure reduction.

34. Ways to intensify homogeneous processes:

A) increase in concentration, decrease in temperature;
B) decrease in pressure and temperature;
C) pressure reduction;
D) the use of catalysts;
E) decrease in concentration.

35. Chemical catalytic processes are called heterogeneous if:

A) reagents, catalyst and products are in one phase;
B) the starting reagents and the catalyst are in the same phase, and the products are in different phases;
C) occur at the phase boundary;
D) only immiscible liquids are components;
E) the use of a catalyst is mandatory.

36. Chemical catalytic processes are called heterogeneous if:

A) reagents, catalyst and products are in one phase;
B) the starting reagents and the catalyst are in the same phase, and the products are in different phases;
C) only immiscible liquids are components;
D) proceed at the interface-reacting substances-catalyst;
E) occur only in the liquid phase.